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Ethanol ( $\text{CH}_3\text{CH}_2\text{OH}$ ) is being used as a gasoline additive or alternative in many parts of the world. (a) Use bond energies to find  $\Delta H^\circ_{\text{rxn}}$  for the combustion of gaseous ethanol. (Assume  $\text{H}_2\text{O}$  forms as a gas.) (b) In its standard state at  $25^\circ\text{C}$ , ethanol is a liquid. Its vaporization requires  $40.5 \text{ kJ/mol}$ . Correct the value from

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