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the empirical formula for a 100.00-g sample of a compound having the following percent composition. a. 94.07% Sulfur and 5.93% hydrogen b. 80.68%

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Chemistry: Supplemental Stoichiometry Problems. Directions: Solve each of the following problems. Assume excess of any reactant that isn't mentioned, unless otherwise specified. Show your work, including proper units, to earn full credit. 1. Given the balanced equation, show what the

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following molar ratios (i.e., the conversion factors) should be.

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aluminum is placed in a 25-mL graduated cylinder containing 10.0 mL of water. The level of water rises to 18.0 mL. Aluminum has a density of 2.7 g/mL. Calculate the mass of the sample.

2. Saturn is about 1 429 000 km from the Sun.

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Answer:  $4.93 \times 10^{-5}$  L or 49.3  $\mu$  L In  
Example 12.2.1 and Example 12.2.2, the  
identity of the limiting reactant has been  
apparent:  $[\text{Au}(\text{CN})_2]^-$ ,  $\text{LaCl}_3$ , ethanol,  
and para -nitrophenol. Stoichiometry

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